

Answer: (a) 1.00 mol (b) 2.00 mol (c) 1.00 mmol (d) 2.00 mmol

هو دصر حود Q2: The air inside a flexible 3.5 L container has a pressure of 115 kPa. What should the volume of the container (25 points) be increased to in order to decrease the pressure to 625 torr? ideal plain Ed an 20

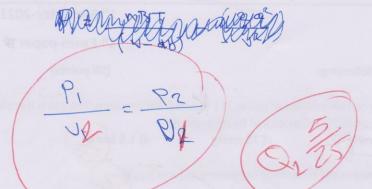
Q3: A 3 dm³ container holds 0.5 moles of N₂ gas at 42 °C. What is the pressure inside the container? (25 points)

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Best wishes

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02



V1= 3-52 V2=? P1=155 KPM P2=625 book

 $P = \frac{NBT}{NBT} \Rightarrow P = \frac{5.5 \times 508 \times 402}{3?} = 0.56 \text{ }$