**Name:**

**Home Work**

**Q-1- (0.2) g of saline solution consisting of mixing (Silver chloride and potassium nitrate) was passed through a cationic exchanger (H+ \_ Form) and the collected solution from the column was calibrated against the sodium hydroxide base ,If the flow volume of the burette was( 20) ml ,**

**calculate :**

**1 . The weight of Silver chloride in the sample ?**

**2. The weight of potassium nitrate in the sample ?**

**3. The Normality concentration of the sodium hydroxide ?**

**4. PPM for potassium nitrate in the solution (5ml) ?**

**5. Write the equations for this question ?**

**6. PH for the come down solution from the column ?**

**If you know the percentage of potassium chloride in the sample is 80 %**

**And the Atomic weight for ( K = 39 , N =14 , O = 16 , Cl = 35.5 , Ag = 108 ,Na = 23**